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Example

\(\backslash\PageIndex{1}\backslash\) How many moles of oxygen react with hydrogen to

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produce 27.6 mol of H₂O? The balanced equation is as follows:
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$.
Solution. Because we are dealing with quantities of H₂O and O₂, we will use a ratio that relates those two substances. Because we are given an amount of H₂O and want to determine an amount of O₂, we will use the ratio that has H₂ ...

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6.4: Mole-Mole Relationships in Chemical Reactions

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Unit 8 Worksheet 1:

Mole relationships For each of the problems below:

a. Write the balanced chemical equation.

b. Identify what is given (with

what is given (with

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units) and what you want to find (with units). c.

Modeling Chemistry Unit 8 Worksheet 1 Mole Relationships

...

6.4 Mole-Mole Relationships in Chemical Reactions. Learning Objective. ... Defend your answer. Explain how a chemical reaction balanced in terms of moles satisfies the law of

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conservation of matter.

Answers: 1 mol NH₃:2 mol O₂:1 mol HNO₃:1 mol H₂O. 3.14 mol.
 $4\text{NH}_3 + 3\text{O}_2 \rightarrow 2\text{N}_2 + 6\text{H}_2\text{O}$; 2.48 mol.

Mole-Mole Relationships in Chemical Reactions

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Mole / Mole Relationships To convert moles of one substance to moles of another we use a ratio obtained from a balanced chemical equation. moles of A \times

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moles of B moles of A

= moles of B For

Example: Suppose we

wanted to determine

how many moles of

oxygen would be

needed to completely

react with 13.3 moles

of C₂H₆: 2C₂H₆ + 7O₂

→ 4CO₂ + 6H₂O ...

Name: Stoichiometry Sheet 1 Mass / Mole Relationships

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2 moles H 2 O 2 x 2.02
g=4.04 g H 2 32.0 g O
2 2 x 18.02 g=36.04 g

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H₂O Figure

(\PageIndex{1}\): This representation of the production of water from oxygen and hydrogen show several ways to interpret the quantitative information of a chemical reaction.

6.4: Mole-Mole Relationships in Chemical Reactions

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Constant. An amount
of substance
containing 6.02×10

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23 particles is called a mole (often abbreviated to mol).; 6.02×10^{23} is called the Avogadro Constant or Avogadro's Number.; The following diagram shows how to convert between Mass, Mole and Number of particles.

Mole, Avogadro Constant & Molar Mass (solutions, examples ...

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