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and your knowledge of periodic trends to answer several examstyle questions like why do atoms get smaller as you move left to right in a period?,

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Use your notes to th answer the following questions, 1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium, 2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum. 3. Why does fluorine have a higher ionization energy than iodine ...

Wasatch County School District / Overview Periodic Properties Worksheet Key 1. State which atom is larger according to periodic trends. a) Br, F Br b) Ca, Mg Ca c) C, Si Si d) He, Ar Ar 2. Explain the difference in size for each of the pairs in Problem 1. The outer electrons of the larger atom are in a subshell further away from the nucleus than the outer

electrons of the smaller atom. 3.

Periodic Properties Worksheet Key

The periodic table is useful for predicting behavior of elements if you know periodic trends in properties. You should be familiar with periodic trends from your prior experience in chemistry. Now that you are an advanced chemistry student, you

will need to explain why they happen and why there are some exceptions.

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